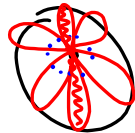


Atomic Orbitals

$n=1$



$n=2$



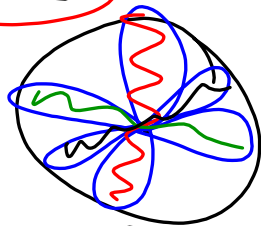
2p (3 of them)

2s (bigger)

sublevels

4 orbitals total

$n=3$



$l=1$

3p (3 of them)

3s

+ 3d orbits (5 of them)

$n=4$ → even more!

Electron Configuration (arrangement of electrons)

- each electron occupies lowest energy orbital available

