

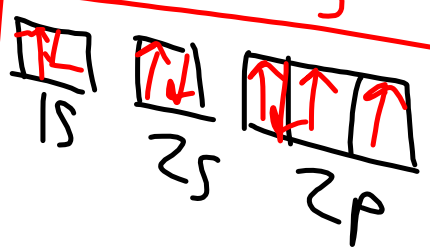
Electron Configurations Con't.

Aufbau principle - each electron occupies the lowest energy level available

Pauli exclusion principle - there are 2 electrons per orbital

• electrons have opposite spins ↑ and ↓ in the same orbital

Hund's Rule: Single electrons with same spin occupy equal energy orbitals before they can occupy same orbital

Element	Atomic #	Orbital Diagram	Electron Configuration
Oxygen	8		$1s^2 2s^2 2p^4$